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18-1 CHAPTER 18 ACID-BASE
EQUILIBRIA Section 18.1: Acids and
Bases in Water Water (H_2O) - the
most important molecule on
earth. Even in pure water, there are
small amounts of ions from the
equilibrium below ("self-ionization of
water" or "auto-ionization).

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model of acids and bases states that an acid contains the element hydrogen and forms ions of this element when it is dissolved in water. a base contains the hydroxide group and dissociates to produce hydroxide ions in aqueous solution

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B solution contains more hydroxide ions than hydrogen ions. False. An aqueous solution of acid causes red litmus paper to turn blue. False. In a typical reaction, if the reacting acid is strong, the base produced is strong.

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Section 118.18.1 Figure 18.1

Rhododendrons flourish in rich, moist soil that is moderately acidic, whereas semper-vivum, commonly called hen and chicks, grows best in drier, slightly basic soil. Introduction to Acids and Bases-!).)DEADifferent models help describe the behavior of acids and bases.

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Section 118.18.1 Figure 18.1

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Acids are named by changing the ending of the alkane to -oic acid. i.e. a 4 carbon acid goes from butane to butanoic acid Acid salts are named by changing the acid ending to -oate i.e the 4 carbon acid goes from butanoic acid to sodium butanoate Esters are named by changing the ending of the alkane to -oate and by using alkyl name i.e. the 4 carbon ester below is methyl butanoate(methyl for OCH₃ portion) O O Anhydrides are dehydrated esters that replace -oic acid with word anhydride Acyl ...

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fertilizers and detergents, but it is also used to make pharmaceuticals,

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