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### Chemical Equilibrium

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At  $450^{\circ}\text{C}$ , the value of the equilibrium constant for the following system is  $6.59 \times 10^{-3}$ . If  $[\text{NH}_3] = 1.23 \times 10^{-4}\text{ M}$  and  $[\text{H}_2] = 2.75 \times 10^{-2}\text{ M}$  at equilibrium, determine the concentration of  $\text{N}_2$  at that point.  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$

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The state of equilibrium can be maintained over time as long as all factors remain the same. The state of equilibrium can be maintained over time as long as all factors change. The chemical...

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a reversible reaction in which there is no longer any change in the concentration of reactants and products. chemical equilibrium. a state of balance in which the rate of a forward reaction equals the rate of the reverse reaction and the concentrations of products and reactants remain unchanged. equilibrium.

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Use the following chemical equation to answer the question below:  $\text{Mg}(\text{s}) + 2\text{H}_2\text{O}(\text{aq}) \rightleftharpoons \text{Mg}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) + \text{H}_2(\text{g})$  If 0.048 g of magnesium completely reacts in 20 s, what is the average reaction rate in moles/second over that time interval? Average rate  $9.9 \times 10^{-5}\text{ mol/s}$  MODERN CHEMISTRY REACTION KINETICS 141 Copyright © by Holt, Rinehart and Winston. All rights reserved.

#### R k H 2 NO 2 3 Use the following chemical equation to ...

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ANS:  $K_p$  (equilibrium constant) is independent of pressure and concentration. 34 .One mole of a compound AB reacts with one mole of a compound CD according to the equation  $\text{AB} + \text{CD} \rightleftharpoons \text{AD} + \text{CB}$ . When equilibrium had been established it was found that 34 mole each of reactant AB and CD had been converted to AD and CB.

#### Chemical Equilibrium Important Questions And Answers

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### ~~Holt Chemistry Concept Review Answers Chapter 4~~

18) The following equilibrium is readily established:  $\text{SO}_2\text{Cl}_2(\text{g}) \rightleftharpoons \text{SO}_2(\text{g}) + \text{Cl}_2(\text{g})$  At equilibrium at 373 K, a 1.00-L reaction vessel contains 0.0106 mol of  $\text{SO}_2\text{Cl}_2$  and 0.0287 mol each of  $\text{SO}_2$  and  $\text{Cl}_2$ . What is  $K_{\text{eq}}$  for the reaction at 373 K? A)12.8 B)2.72 C)0.0781 D)2.39 E)0.418

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